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ENTHALPIES OF FORMATION OF WF_5N_3 , $ReF_5(NCl)$ AND OF OTHER
TRANSITION METAL FLUORIDES WITH METAL-NITROGEN BONDS

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From alkaline hydrolysis measurements the enthalpies of formation of $WF_5N_3(s)$ and of $ReF_5(NCl)(s)$ have been estimated to be -1170 and -1195 kJ mol^{-1} respectively. Assuming the metal-fluoride bond strength to be the same as in the parent fluorides, $D(W-N)$ and $D(Re=N)$ are found to be about 323 and 680 kJ mol^{-1} respectively. The latter value compares with the $D(Re=O)$ in $ReOF_5$ of 635 kJ mol^{-1} . Similar data will be presented for WF_4NH and for $MoF_4(NCl)$.

We wish to thank Professor K. Dehnicke for a sample of $MoF_4(NCl)$.